



Cite this: *Phys. Chem. Chem. Phys.*,
2022, 24, 6185

Photoswitching activation of a ferrocenyl-stilbene analogue by its covalent grafting to gold†

Francesc Bejarano,^a Diego Gutiérrez,^{ab} José Catalán-Toledo,^{ID a}
Daniel Roca-Sanjuán,^{ID c} Johannes Gierschner,^{ID *d} Jaume Veciana,^{ID a}
Marta Mas-Torrent,^{ID a} Concepció Rovira,^{ID a} and Núria Crivillers^{ID *a}

Until now, surface-deposited stilbenes have been much less studied than other photochromic systems. Here, an asymmetrically substituted styrene incorporating a redox-active ferrocene moiety and a terminal alkyne group has been synthesised to investigate its photoisomerization in solution, and upon the formation of chemisorbed self-assembled monolayers through a carbon–gold bond formation. Charge transport measurements across the monolayers reveal that upon chemical linkage to the gold substrate there is an alteration of the isomerization pathway, which favours the *trans* to *cis* conversion, which is not observed in solution. The experimental observations are interpreted based on quantum chemistry calculations.

Received 3rd November 2021,
Accepted 21st February 2022

DOI: 10.1039/d1cp05012e

rsc.li/pccp

1. Introduction

In the field of molecular electronics,^{1–3} there is a continuous search to attain stable, reproducible and efficient metal adsorbed functional molecules displaying amplification, rectification or information storage functions.⁴ With this aim, surface confined molecular switches⁵ have been extensively explored,^{6–8} and with their implementation into molecular-based devices have been applied as memories,^{9,10} conductance switches^{11–13} and, to tune surface properties such as the metal work function,^{14,15} surface wettability^{16–18} and the substrate optical and magnetic properties.¹⁹ Among the possible external stimuli required to induce the switching process, the remote, selective control and spatial and temporal precision, that light irradiation provides, being additionally non-invasive and clean, makes it a very appealing toggling source for devices fabrication.^{20–27} Stilbenes, azobenzenes, diarylethenes and spiropyranes have been widely studied photochromic systems.^{28–31} It is well known that the surface grafting of a molecular switch can strongly influence the switching function.^{32–34} The photoisomerization of

stilbenes on surfaces has been much less exploited than that of azobenzenes and diarylethenes.^{35–38} The reason for this may lie in the possibility of subsequent photochemical reactions, *i.e.*, phenanthrene formation and cyclodimerization reactions, which can lower the *cis-trans* photoisomerization yield.

The herein reported work aims to fill this gap of research by studying the light response of an asymmetrically substituted styrene SAM incorporating an electrochemically active moiety, *i.e.* a photo- and redox-active system (compound **1** in Fig. 1).^{39,40} **1** was chosen as the target system for two main reasons: (i) to employ a terminal alkyne anchoring group which spontaneously reacts with the electrode through a C–Au bond establishing a robust and electronically effective contact with the metal substrate and, (ii) to incorporate as a styrene substituent a 5-membered ring which could prevent possible phenanthrene generation upon illumination of the SAM. For this, a ferrocenyl moiety was selected since additionally its redox-ability allows its electrochemical identification. In this work, it is shown that in solution, *trans* to *cis* isomerization is inefficient, but becomes favoured upon covalently anchoring **1** on the Au surface. This effect was revealed by charge transport measurements across the

^a Institut de Ciència de Materials de Barcelona (ICMAB-CSIC) and CIBER-BBN, Campus de la UAB, 08193 Bellaterra, Spain. E-mail: ncrivillers@icmab.es

^b Leitat Technological Center (LEITAT), Carrer Innovació, 2, 08225 Terrassa, Spain

^c Institute of Molecular Science, University of Valencia, P.O. Box 22085, ES-46071 Valencia, Spain

^d Madrid Institute for Advanced Studies, IMDEA Nanoscience, Calle Faraday 9, Campus Cantoblanco, 28049 Madrid, Spain.

E-mail: johannes.gierschner@imdea.org

† Electronic supplementary information (ESI) available: Synthesis and characterization of **1**, additional, in solution and on SAM, characterization and further theoretical details. See DOI: 10.1039/d1cp05012e

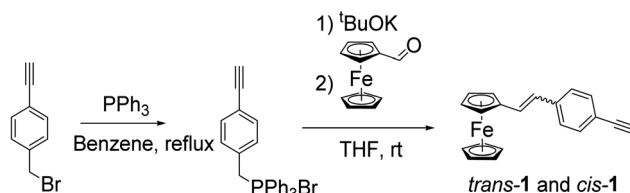


Fig. 1 Synthesis of the ferrocenyl-stilbene analogue **1**.



SAMs and was supported by a theoretically determined barrier energy interconversion between the *trans* and *cis* forms for the **1-Au**.

2. Experimental and computational details

2.1. SAM preparation

All the glassware used for the preparation of SAMs was first cleaned following standard procedures and then further washed by overnight immersion in a 2% v/v Hellmanex II solution in MilliQ water, thoroughly rinsed with MilliQ water and dried in an oven at 80 °C overnight. When Teflon material was used, it was cleaned with piranha solution at 100 °C and carefully rinsed with MilliQ water as well. Ultra-smooth template-stripped gold and silver substrates (Au^{TS}) were prepared following a reported procedure.⁴¹ The thickness of the gold layer was 200 nm. A 1.2 mM solution of compound **trans-1** or **cis-1** in freshly distilled toluene were dissolved under argon atmosphere in a flat-bottom flask. Then, the Au^{TS} substrates were immersed in the solution and the flask was closed and sealed. Three vacuum/Ar cycles were carried out to remove the residual atmospheric air. After 24 hours at room temperature and in dark, the substrates were removed from the flask and thoroughly rinsed with toluene and dried in a nitrogen stream.

2.2. Charge transport measurements

The Gallium–Indium eutectic (EGaIn) was used as the top-electrode. An electrometer Keithley 2604B was used to perform the electrical measurements. The current density *versus* potential curves (*J/V*) were obtained with an integration time settle between 50 ms to 250 ms depending on the current value and a waiting time between two consecutive potential values of typically 100 ms (100 mV/100 ms). The EGaIn was purchased from Sigma-Aldrich ($\geq 99.99\%$ trace metals basis). The top-electrode was biased and the bottom electrode was grounded. EGaIn/Ga₂O₃/ferrocenyl-stilbene/substrate junctions were fabricated by placing in contact a EGaIn/Ga₂O₃ drop hanging from a Hamilton syringe. Different junctions were prepared and several traces for each junction were registered in the form of *J/V* curves. For each set of measurements (*i.e.* each junction) at a given point of the sample, the freshly prepared EGaIn/GaO_x cone-shaped was placed in contact with the SAM/Au surface, with a geometrical contact diameter of around 20 μm .^{42,43} In order to obtain representative values from the *J/V* curves, the data was treated statistically to determine the mean, the standard deviation and the 95% confidence intervals. For each potential value (voltage steps of 0.1 V) the data were plotted in histograms in $\log 10|J|$ and a Gaussian fit was applied.

2.3. Calculations

(Time-dependent) density functional calculations, (TD)-DFT, were employed using the B3LYP as well as the long-range corrected CAM-B3LYP functional and the 6-311G* basis set (LANL2DZ for Fe, Au) as defined in the Gaussian program

package.⁴⁴ Static dipole moments were extracted from the ground state DFT calculations. Vertical energies were obtained as single point TD-DFT calculations on the DFT-optimized ground state geometries of the *trans* and *cis* isomers. Adiabatic energies were obtained from TD-DFT geometry optimizations of the respective first singlet and triplet excited states (*S*₁, *T*₁).

3. Results and discussion

3.1. Synthesis and SAM preparation and characterization

The synthesis of compound **1** was achieved following the synthetic route shown in Fig. 1. An isomeric mixture was obtained from a Wittig reaction between the triphenylphosphonium bromide derivative and the ferrocene carboxaldehyde. Small fractions of pure **trans-1** and **cis-1** were isolated, along with a large fraction of the diastereomeric mixture that was quantitatively converted to the *trans* form in the presence of catalytic iodine upon toluene reflux (see ESI† for details on the synthesis and characterization).

The cyclic voltammetry (CV) of both isomers gave very similar redox potential, $E^{1/2} = 0.44$ V and $E^{1/2} = 0.46$ V (*vs.* Ag/AgCl) for the **trans-1** and **cis-1**, respectively. UV-Vis absorption spectroscopy showed two clearly different spectra for both isomers (Fig. 2) with a significant absorption enhancement (hyperchromic effect) and a bathochromic (red) shift for the main absorption band *A*₃ of **trans-1** *vs.* **cis-1**. Both effects are ascribed to enlarged π -conjugation for the *trans* form compared to the sterically restricted non coplanar *cis* configuration. This is indeed shown by DFT geometry optimization (Fig. S7, ESI†), giving a dihedral angle in the vinylene unit of 180° for **trans-1**, and 6° for **cis-1**. Furthermore, our TD-DFT based natural transition orbital (NTO) analysis in Fig. S11 (ESI†) indeed ascribe *A*₃ to an extended locally excited (LE) $\pi\pi^*$ -type state (mainly *S*₉); the extension of the π -system by the acetylene moiety is also responsible for a considerable bathochromic shift of *A*₃ *vs.* *trans*-stilbene.

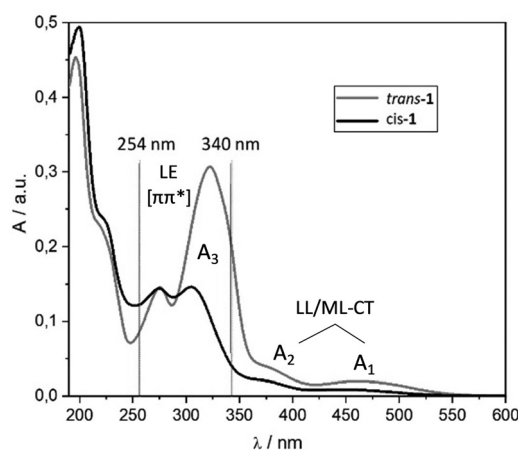


Fig. 2 Absorption spectra of **trans-1** and **cis-1** in acetonitrile. Straight lines correspond to the irradiation wavelengths used in assayed isomerizations. Indicated transitions: LE = locally excited state; LL = ligand-to-ligand; ML = metal-to-ligand, CT = charge transfer.



The two weak absorption bands $A_{1,2}$ in Fig. 2 are assigned by the NTO analysis to a number of mixed ligand/metal-to-ligand charge-transfer (CT) transitions, formed within the ferrocene moiety with some participation of the adjunct styryl unit; for the NTOs of the first excited state (S_1) in **trans-1**, see Fig. 3. The weak red-shift of $A_{1,2}$ for **trans-1** vs. **cis-1** is again attributed to enlarged conjugation for the *trans* form; the shift is however not as pronounced as for the LE state (A_3).

The photoisomerization studies were first carried out in solution. The **cis-1** $\rightarrow$ **trans-1** conversion was demonstrated to happen by the intensity increasing and red-shifting of the band at 305 nm. This process was induced by irradiating an acetonitrile solution (0.05 mM) of **cis-1** at $\lambda = 254$ nm for 100 min (see ESI[†], Fig. S1), achieving approximately 76% of isomerization. On the contrary, the **trans-1** $\rightarrow$ **cis-1** process was not accomplished by irradiating at $\lambda = 366$ and 340 nm at several concentrations and solvents (see ESI[†], Fig. S2). The observed one-way isomerization was previously reported for the parent styryl ferrocene systems (**StFc**, *i.e.* without the acetylene moiety).⁴⁵ It was suggested that the presence of ferrocene gives rise to an energetically accessible triplet state, which permits a one-way triplet state isomerization pathway.⁴⁵ Our NTO analysis of **trans-1** in Fig. 3 indeed indicates significant participation of Fe d-orbitals in the S_1 state to promote inter-system crossing (ISC) to the triplet manifold by effective spin-orbit coupling (SOC).

Bearing in mind the known influence of the surface on the isomerization pathway of photochromic molecules adsorbed on surfaces,³³ the study of the irradiation response of the SAMs based on **cis-1** and **trans-1** was pursued. First, SAMs of compound **cis-1** and **trans-1** were separately prepared from solution on gold template-stripped (Au^{TS})⁴¹ by the spontaneous formation of the C-Au bond^{46–50} (see ESI[†] for details on the SAM formation). Both SAMs were successfully characterized by CV and X-Ray photoelectron spectroscopy (XPS) (see ESI[†], Fig. S3 and S4, respectively). In the CV, the ratio between the maximum intensity of both reduction and oxidation peaks gave values of 1.25 (**trans-1-SAM**) and 1.19 (**cis-1-SAM**) that are very

close to 1, indicating a high reversibility of the redox processes. The peak splitting ($\Delta E = E_{\text{ox}} - E_{\text{red}}$) at 100 mV s^{-1} was 40 mV for **trans-1-SAM** and 30 mV for **cis-1-SAM**. These differences between both SAMs could be attributed to certain different packing structures of the two isomers within the layer. Such small ΔE value is also in agreement with an adsorbed redox reversible system, in which diffusion does not play a role. Furthermore, from the integration of the redox peak is possible to determine the molecular surface coverage (see ESI[†]), which shows similar values, $1.8 \pm 0.2 \times 10^{-10}$ and $1.6 \pm 0.2 \times 10^{-10} \text{ mol cm}^{-2}$ for **trans-1-SAM** and **cis-1-SAM**, respectively. In the XPS spectra (Fig. S4, ESI[†]), the presence of iron on the surface was confirmed by the two peaks at 707.5 (2p_{3/2}) and 720.3 (2p_{1/2}) eV with the expected intensity ratio (1.8). An insignificant peak in the C1s spectrum around 288.6 eV (attributable to C=O) pointed to a negligible molecule oxidation during the grafting process.

3.2. Charge transport measurements

For the isomerization study of the SAMs, charge transport measurements across the layers were performed using a eutectic gallium–indium alloy (EGaIn) tip as the top electrode to contact the SAM/Au substrate forming the desired molecular junction. The liquid metal permits to assess the charge transport properties across molecular monolayers, by electrically contacting them without damaging the organic layer.⁴³ In contrast to the extensive literature related to non-functional (*e.g.* alkanethiolated SAMs⁵¹) or redox-based SAMs,^{52–55} the use of the EGaIn technique to study photochromic monolayers is scarce.^{56,57} First, **trans**- and **cis**-1-SAMs were independently investigated. In these measurements, the surface was grounded and the applied voltage was swept in a potential window of ± 0.5 V. To obtain significant statistics, the current density *vs.* voltage (*J*/*V*) curves were acquired at different points on the sample, and this process was done for several SAMs (a detailed statistic and measurement protocol is described in the ESI[†]). Fig. 4a shows a scheme of the EGaIn/ Ga_2O_3 /**trans-1-SAM**/Au junction. The yield of successful junctions formed was around 80% for both SAMs. Interestingly, the current density at +0.5 V was about 1.3 orders of magnitude higher for the *cis*-based SAM

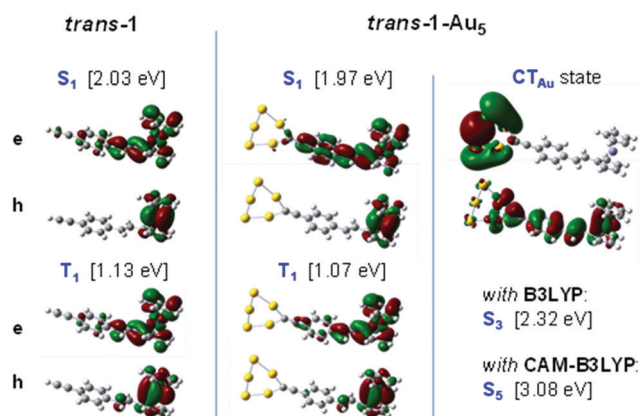


Fig. 3 Natural transition orbitals (NTOs; with h = hole, e = electron) for the S_1 and T_1 states of **trans-1** and **trans-1-Au₅**, as well as the CT_{Au} state, as calculated by TD-DFT (B3LYP); numbers in brackets indicate the vertical transition energies.

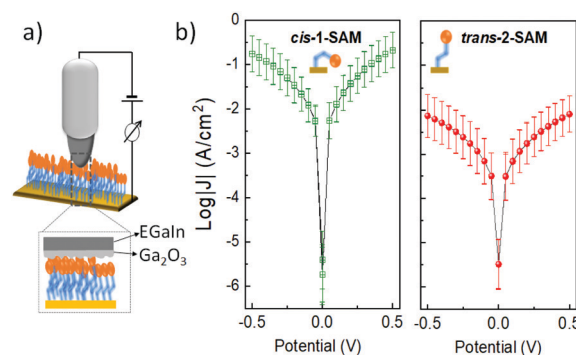


Fig. 4 (a) Scheme of the EGaIn/ Ga_2O_3 /**trans-1-SAM**/Au. (b) *J*/*V* plots acquired for **trans-1-SAM** and **cis-1-SAM**. Dots (circles and squares) indicate mean values while the bars stand for the standard deviation.



than for the *trans* one (Fig. 4b, see heatmaps in Fig. S5a and S6a, ESI†). The cause for this difference could be attributed to two factors: (i) the length difference between the two configurations (*cis* and *trans*). In tunnelling regime, the measured current through the monolayer decreases exponentially with distance. The shorter length of the *cis* isomer results in shorter molecular junctions that might exhibit smaller tunnelling barriers^{51,58} as observed for *cis/trans* azobenzene-based SAMs^{20,59} and, (ii) certain contribution of a different molecular dipole. In large molecular junctions, the influence of the molecular dipole on the measured current is not easy to correlate due the complexity of the system (*i.e.* depolarization effects, monolayer disorder, *etc.*) but, it has been observed that the dipole magnitude and orientation towards the surface normal can alter relevant parameters such as the molecule-electrode coupling strength and/or surface work function, which may directly impact on the resulting rate of charge transport.⁶⁰ From the DFT-optimized geometries of the isolated systems, a small molecular length difference was calculated for both isomers (1.1 Å measuring the longest distance possible, *i.e.* from the acetylene-H to the terminal of the non-functionalized cyclopentadiene-unit of the ferrocene, Fig. S7, ESI†). Additionally, different magnitude and orientation of the molecular dipole was calculated, being 1.00 D and 1.59 D for the *trans* and *cis* configurations, respectively (Fig. S8, ESI†). These given values can vary upon the adsorption of the molecules on the surface together with additional dipole-induced effects^{60–62} but, it is clearly expected that both parameters are important in determining the different current responses observed.

These results undoubtedly demonstrated that the electrical response could be used as a reading output to differentiate the two grafted isomers. Hence, the “EGaIn technique” was employed to investigate the impact of the molecule–gold bond formation on the light-driven *cis-trans* olefin isomerization in SAMs composed of **1**. For that, SAMs of pure *trans* and *cis* isomers were illuminated at appropriate wavelengths. First, the *trans-1* → *cis-1* was studied once the molecules were covalently linked to the Au surface. For doing that, a *trans-1-SAM* was *ex situ* irradiated at $\lambda = 340$ nm for 40 minutes (see ESI† for more details). The *JV* curves were acquired before and after the irradiation. Fig. 5 depicts the histograms showing the density distribution of $\log J$ at +0.5V, before (Fig. 5a) and after the SAM illumination (Fig. 4b) (see the corresponding heatmaps in the ESI† Fig. S5a and b). Strikingly, while data before irradiation can be fitted to a Gaussian distribution with a maximum corresponding to a pure *trans-1-SAM* ($\log J \approx 2.1$), after the irradiation, a larger population around $\log J \approx 1.2$ appears, which is attributed to the *cis* isomer since the maximum is very close to the one shown by the SAM prepared from pure *cis-1* (Fig. 5c). In the histograms, counts for each $\log J$ were divided by its bar area in order to roughly approximate the Gaussian distribution areas to the *trans-1/cis-1* isomer ratio. By analysing the Gaussian areas, the conversion was estimated to occur in a *c.a.* 82% yield. To support the molecular configuration change, as the reason of the different observed *J* values, the same

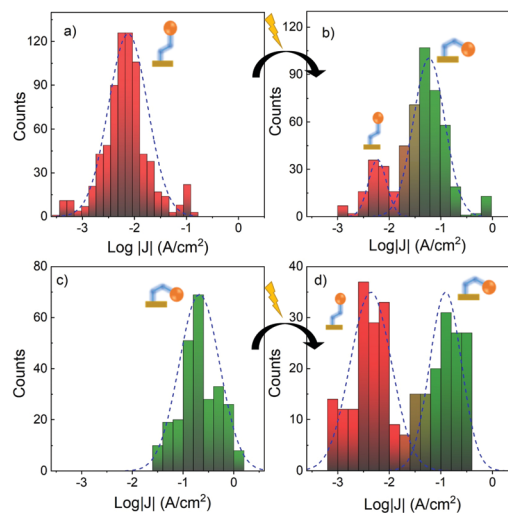


Fig. 5 Histogram showing the density distribution of $\log|J|$ at +0.5 V for (a) *trans-1-SAM* before and (b) after irradiation for 40 minutes at $\lambda = 340$ nm; (c) *cis-1-SAM* before and (d) after irradiation for 40 minutes at $\lambda = 254$ nm.

experiment was carried out irradiating at $\lambda = 365$ nm. In this case only a partial isomerization was observed (Fig. S9, ESI†) which, based on the results in solution, is in accordance with a lower absorption of compound *trans-1* and importantly, a smaller difference between *trans* and *cis* absorption at this wave length. It may, however, also point to a need for excess energy to promote effective isomerization (*vide infra*). The resulting SAM was further irradiated with the 240 nm lamp to induce the back isomerization (*i.e.* *trans* → *cis* → *trans*) and, although the reversibility yield is low, the heatmap plots show that the population assigned to the *cis* isomer diminishes (Fig. S9, ESI†).

The *cis-1-SAM* → *trans-1-SAM* conversion was also investigated starting from a prepared *cis-1-SAM*. The tunnelling junctions were measured before and after *ex situ* illumination at $\lambda = 254$ nm and, histograms of the $\log|J|$ for both measurements, before and after illumination, were compared (Fig. 5c and d). Initially, the data acquired for the pure *cis-1-SAM* can be fitted to a Gaussian distribution with a maximum at $\log|J| \approx 1$ and the histogram corresponding to the illuminated sample shows two populations that can be separately adjusted to two Gaussian distributions, showing maximums around -0.8 and -2.3 . These values are very close to those measured in dark for pure *cis-1* and *trans-1-SAMs*, respectively, thus pointing to a partial isomerization, which is also very clear in the corresponding heatmaps (Fig. S6, ESI†). Under irradiation, the higher-conductance population diminishes and the lower-conductance population appears. An approximated value of 56% of conversion yield can be deduced upon area integration under both Gaussian curves. A plausible explanation for this partial switching could be the fact of having an inhomogeneous SAM organization, (*i.e.* islands with different packing motives), some offering major geometrical constrain for the isomerization to take place. Hence, our observations do not correspond



to single-molecule effects. Therefore, intermolecular interactions and defect boundaries may be playing a role on the isomerization rate.⁶³ These different factors may promote to an uneven spatial distribution of the two isomers on the surface upon irradiation.

Control experiments were also performed with a ferrocene-based SAM but without incorporating the photoactive unit. For that, 6-(ferrocenyl)hexanethiol (FcC₆SH) was employed to prepare SAMs on Au^{TS} (Fig. S10, ESI†). The measured current for the EGaIn/Ga₂O₃/FcC₆S/Au^{TS} junctions before and after the irradiation at different wavelengths showed insignificant changes. These results evidence that the measured current changes upon the irradiation of *trans*- and *cis*-1-SAMs are attributable to the carbon-carbon double bond isomerization. That is, by contrast with the experiments in solution, the *trans*-1 isomerization seems to take place in the SAM.

3.3. Theoretical analysis

To interpret the experimentally observed activation of two-way isomerization on the Au surface *vs.* one-way isomerization in solution, we investigated the isomerization by theoretical means. The current state-of-the-art understanding of photoisomerization in stilbenoid compounds and polyene molecules involves a crossing between electronic states, either by a conical intersection (CI) and/or a singlet/triplet crossing (STC),^{64–68} depending on the molecular structure.

For the parent styryl-ferrocene (StFc) compound in solution, an (inefficient) one-way *cis*-photo-isomerization path *via* the T₁ state was proposed, *i.e.* ³c* → ¹t through a STC.⁴³ This is adapted for compound **1** in solution (Fig. 6) due to the pronounced structural similarity, and the similar isomerization behaviour in solution. In fact, the TD-DFT calculations give very similar T₁ and S₁ energies for **1** and StFc in both isomers, see Table S1 (ESI†). Furthermore, the vinyl double bond length *d*_{DB} does not change upon introduction of the acetylene unit; and remains essentially unaltered (1.347 ± 0.002 Å) when going from S₀ to T₁ and S₁ (Table S1, ESI†). We emphasize that this is fundamentally different to stilbene, where *d*_{DB} is strongly

labilized in S₁,⁶⁹ which opens the singlet isomerization path for the *trans* and *cis* excited states (¹t*, ¹c*).⁶⁸ On the other hand, in **1** and StFc, the singlet path is ineffective for this reason. Instead, the (one-way) triplet path *via* ³c* becomes operative by ISC (see the blue arrows in the deactivation scheme of Fig. 6). The T₁ path is however not very effective despite the long T₁ lifetime, due to the negligible labilization of *d*_{DB} in T₁; the STC can therefore only be reached from ³c*, but not from ³t* which is found ~0.2 eV below ³c* (see Table S1, ESI†).

Upon grafting to Au, the geometry in the S₀ and T₁ states of **1-Au** remains unchanged *vs.* **1**, see *d*_{DB} in Table S1 (ESI†); this gives little evidence for a change to a two-way triplet path isomerization on gold. A different scenario is therefore proposed for the Au-grafted species. In fact, TD-DFT predicts the appearance of a low-lying CT singlet state, where the electron (e), corresponding to the unoccupied NTO, is localized in the Au moiety; this CT_{Au} state is associated with the high electron affinity (EA) of Au. Using the B3LYP functional, CT_{Au} becomes S₁ in **1-Au**, and S₃ in **1-Au₅**, see Fig. 3; a similar result is obtained by CAM-B3LYP (giving S₅).^{69–71} The CT_{Au} state is characterized by a very significant elongation of *d*_{DB}; in fact, *d*_{DB} = 1.372 Å is calculated for *trans*-**1-Au** (1.376 Å for *cis*-**1-Au**), being much longer than in **1** (1.349 Å). Such double bond labilization is expected to effectively open a path to a CI;⁷² we infer that the CI can be effectively accessed from a higher excited singlet state by providing sufficient excess energy.

The stabilizing effect of the substituent's EA on the CI has been shown in distinct theoretical studies for asymmetrically substituted ethylenes; in fact, substituents with large EA differences (and in particular with opposite sign) at each side of the ethylene bond stabilize a zwitterionic structure which in all lead to an energetic lowering of the CI.^{72–75} As the size of the system is too large to perform multi-reference quantum-chemical calculations to explore the non-radiative decay path and to characterize the crossing, we propose to utilize a qualitative picture, estimating the deactivation path from the S₁ energies and the relative EA of the substituents;⁷⁶ data, which are accessible by simple and affordable (TD)DFT calculations. Such modelling approach has been shown to be valid in stilbenoids.⁷²

In **1** (as well as in StFc), the EA is negative like in ferrocene (Fc) and of similar size, giving rise to a high-lying CI on the S₁ potential energy surface,⁷⁶ so that ISC is the dominant path. From here, the STC can only be effectively reached from the higher lying ³c*, while there is a substantial barrier from the low energetic ³t*. On the contrary, Au-substitution dramatically increases the EA from −0.65 eV to 1.49 eV (Au₁; 1.86 eV for Au₅), due to the good electron acceptor property of Au, which should strongly stabilize the CI. Therefore, an essentially barrierless path to the CI is now accessible in **1-Au_n** both from ¹t* and ¹c*, successfully competing with the ISC path to the triplet manifold. This semi-quantitative TD-DFT-based scenario gives thus a reasonable, but nevertheless preliminary interpretation of the experimentally observed two-way isomerization for **1-Au_n** *vs.* the one-way isomerization of **1**; further assessment of the current

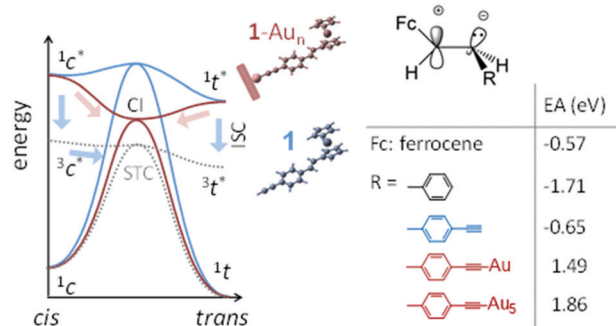


Fig. 6 Potential energy surface of **1** (blue) and **1-Au** (red) for the electronic ground state (¹c, ¹t) and singlet and triplet excited state (¹c*, ¹t* and ³c*, ³t*, respectively) along the main isomerization coordinate (double bond twist), depicting the deactivation pathways (blue and red arrows). The sketch depicts the zwitterionic structure at the state crossing point (STC). Electron affinities (EA) of the different substituents are given in the table.



hypothesis by a closer inspection of the CI/STC region will be subject of future investigations.

4. Conclusions

An asymmetrically substituted styrene incorporating an electrochemically active ferrocene moiety has been synthesized. The photoisomerization studies performed in solution and on surface have revealed a high influence of the presence of Au on the efficiency of the isomerization. While the *trans* to *cis* isomerization is inefficient in solution, it is experimentally demonstrated and theoretically interpreted that this becomes favoured upon the anchoring of the photoactive molecule onto the Au surface. Charge transport measurements across EGaIn/Ga₂O₃/SAM/Au^{TS} junctions have been employed to demonstrate this effect. The combined experimental and theoretical data suggest a change of the photo-isomerization pathway from a one-way triplet path in solution to a two-way singlet path upon grafting on gold due to Au's high electronegativity. Hence, the modulation of the electronic structure of the less exploited photochromic styrenes, which *a priori* have poor photoresponse, by means of their anchor to a metal can be a synthetic strategy to increase their isomerization efficiency and thus, their applicability.

Author contributions

F. B., D. G. and J. C.-T. performed the experimental investigation (synthesis, SAM preparation and electrical characterization) and D. R.-S. and J. G. conducted the theoretical analysis. C. R. and N. C. supervised the work. J. G. and N. C. wrote the original draft; all authors have discussed the results and have contributed to review and edit the final manuscript.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was funded by the Spanish Ministry project GENESIS PID2019-111682RB-I00, the Generalitat de Catalunya (2017-SGR-918) and the Spanish Ministry of Economy and Competitiveness, through the "Severo Ochoa" Programme for Centers of Excellence in R&D (FUNFUTURE CEX2019-000917-S and CIBER-BBN. The work in Valencia and Madrid is supported by the MINECO-FEDER project CTQ2017-87054. The work in Madrid is further supported by the Spanish Science Ministry through the SEV-2016-0686 and CEX2020-001039-S grants and by the Campus of International Excellence (CEI) UAM + CSIC. D. R.-S. acknowledges the MINECO Ramón y Cajal grant (RYC-2015-19234) and the MDM-2015-0538 María de Maeztu Programme. J. C.-T. thanks (DOC-FAM fellowship, grant agreement Nr.754397 through H2020-MSCA-COFUND-2016). F. B. and J. C.-T. are enrolled in the Materials Science PhD program

of UAB. F. B. thanks the Spanish Ministry for the FPU fellowship.

References

- 1 A. Vilan, D. Aswal and D. Cahen, Large-Area, *Chem. Rev.*, 2017, **117**, 4248–4286.
- 2 D. Xiang, X. Wang, C. Jia, T. Lee and X. Guo, *Chem. Rev.*, 2016, **116**, 4318–4440.
- 3 R. L. Carroll and C. B. Gorman, *Angew. Chem., Int. Ed.*, 2002, **41**, 4378–4400.
- 4 C. Joachim, J. K. Gimzewski and A. Aviram, *Nature*, 2000, **408**, 541–548.
- 5 J. D. Harris, M. J. Moran and I. Aprahamian, *Proc. Natl. Acad. Sci. U. S. A.*, 2018, **115**, 9414–9422.
- 6 N. Katsonis, M. Lubomska, M. M. Pollard, B. L. Feringa and P. Rudolf, *Prog. Surf. Sci.*, 2007, **82**, 407–434.
- 7 M. Weinelt and F. von Oppen, *J. Phys.: Condens. Matter*, 2012, **24**, 390201.
- 8 W. R. Browne and B. L. Feringa, *Annu. Rev. Phys. Chem.*, 2009, **60**, 407–428.
- 9 M. Mas-Torrent, C. Rovira and J. Veciana, *Adv. Mater.*, 2013, **25**, 462.
- 10 H. Zhu and Q. Li, *Appl. Sci.*, 2016, **6**, 7–21.
- 11 E. Marchante, N. Crivillers, M. Buhl, J. Veciana and M. Mas-Torrent, *Angew. Chem., Int. Ed.*, 2016, **55**, 368–372.
- 12 G. Reece, C. Lotze, D. Sysoiev, T. Huhn and K. J. Franke, *ACS Nano*, 2016, **10**, 10555–10562.
- 13 S. Jan Van Der Molen and P. Liljeroth, *J. Phys.: Condens. Matter*, 2010, **22**, 133001–133031.
- 14 V. Diez-Cabanes, A. Gómez, M. Souto, N. González-Pato, J. Cornil, J. Veciana and I. Ratera, *J. Mater. Chem. C*, 2019, **7**, 7418–7426.
- 15 N. Crivillers, S. Osella, C. Van Dyck, G. M. Lazzerini, D. Cornil, A. Liscio, F. Di Stasio, S. Mian, O. Fenwick, F. Reinders, M. Neuburger, E. Treossi, M. Mayor, V. Palermo, F. Cacialli, J. Cornil and P. Samorì, *Adv. Mater.*, 2013, **25**, 432–436.
- 16 K.-Y. Chen, G. T. C. Oleksii Ivashenko, J. Robertus, G. Kistemaker, J. C. M. London, W. R. Browne, P. Rudolf and B. L. Feringa, *J. Am. Chem. Soc.*, 2014, **136**, 3219–3224.
- 17 C. Simão, M. Mas-Torrent, J. Veciana and C. Rovira, *Nano Lett.*, 2011, **11**, 4382–4385.
- 18 R. Rosario, D. Gust, M. Hayes, F. Jahnke, J. Springer and A. A. Garcia, *Langmuir*, 2002, **18**, 8062–8069.
- 19 C. Simão, M. Mas-Torrent, N. Crivillers, V. Lloveras, J. M. Artés, P. Gorostiza, J. Veciana and C. Rovira, *Nat. Chem.*, 2011, **3**, 359–364.
- 20 V. Ferri, M. Elbing, G. Pace, M. D. Dickey, M. Zharnikov, P. Samor, M. Mayor and M. A. Rampi, *Angew. Chem., Int. Ed.*, 2008, **47**, 3407–3409.
- 21 C. Jia, A. Migliore, N. Xin, S. Huang, J. Wang, Q. Yang, S. Wang, H. Chen, D. Wang, B. Feng, Z. Liu, G. Zhang, D. H. Qu, H. Tian, M. A. Ratner, H. Q. Xu, A. Nitzan and X. Guo, *Science*, 2016, **352**, 1443–1445.



- 22 D. Dulić, S. J. van der Molen, T. Kudernac, H. T. Jonkman, J. J. D. de Jong, T. N. Bowden, J. van Esch, B. L. Feringa and B. J. van Wees, *Phys. Rev. Lett.*, 2003, **91**, 207402–207405.
- 23 E. Orgiu and P. Samorì, *Adv. Mater.*, 2014, **26**, 1827–1844.
- 24 P. Pourhossein, R. K. Vijayaraghavan, S. C. J. Meskers and R. C. Chiechi, *Nat. Commun.*, 2016, **7**, 11749.
- 25 K. Smaali, S. Lenfant, S. Karpe, M. Oçafrain, P. Blanchard, D. Deresmes, S. Godey, A. Rochefort, J. Roncali and D. Vuillaume, *ACS Nano*, 2010, **4**, 2411–2421.
- 26 L. Wang and Q. Li, *Chem. Soc. Rev.*, 2018, **47**, 1044–1097.
- 27 X. Huang and T. Li, *J. Mater. Chem. C*, 2020, **8**, 821–848.
- 28 D. Bléger and S. Hecht, *Angew. Chem., Int. Ed.*, 2015, **54**, 11338–11349.
- 29 H. M. D. Bandara and S. C. Burdette, *Chem. Soc. Rev.*, 2012, **41**, 1809–1825.
- 30 H. Spreitzer, M. Scholz, G. Gescheidt and J. Daub, *Liebigs Ann.*, 1996, 2069–2077.
- 31 J. Zhang, Q. Zou and H. Tian, *Adv. Mater.*, 2013, **25**, 378–399.
- 32 S. Jaekel, A. Richter, R. Lindner, R. Bechstein, C. Nacci, S. Hecht, A. Kühnle and L. Grill, *ACS Nano*, 2018, **12**, 1821–1828.
- 33 P. Tegeder, *J. Phys.: Condens. Matter*, 2012, **24**, 394001.
- 34 S. Hagen, P. Kate, F. Leyssner, D. Nandi, M. Wolf and P. Tegeder, *J. Chem. Phys.*, 2008, **129**, 164102–164109.
- 35 J. Zhang, J. K. Whitesell and M. A. Fox, *J. Phys. Chem. B*, 2003, **107**, 6051–6055.
- 36 J. Hu, J. Zhang, F. Liu, K. Kittredge, J. K. Whitesell and M. A. Fox, *J. Am. Chem. Soc.*, 2001, **123**, 1464–1470.
- 37 C. S. Tsai, J. K. Wang, R. T. Skodje and J. C. Lin, *J. Am. Chem. Soc.*, 2005, **127**, 10788–10789.
- 38 M. O. Wolf and M. A. Fox, *Langmuir*, 2002, **12**, 955–962.
- 39 J. G. Rodríguez, M. Gayo and I. Fonseca, *J. Organomet. Chem.*, 1997, **534**, 35–42.
- 40 N. Tsuboya, R. Hamasaki, M. Ito, M. Mitsuishi, T. Miyashita and Y. Yamamoto, *J. Mater. Chem.*, 2003, **13**, 511–513.
- 41 E. A. Weiss, G. K. Kaufman, J. K. Kriebel, Z. Li, R. Schalek and G. M. Whitesides, *Langmuir*, 2007, **23**, 9686–9694.
- 42 X. Chen, H. Hu, J. Trasobares and C. A. Nijhuis, *ACS Appl. Mater. Interfaces*, 2019, **11**, 21018–21029.
- 43 R. C. Chiechi, E. A. Weiss, M. D. Dickey and G. M. Whitesides, *Angew. Chem., Int. Ed.*, 2008, **47**, 142–144.
- 44 M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Menucci and G. A. Petersson, *Gaussian 09, Revis. D.01*, Gaussian, Inc., Wallingford, CT, 2009.
- 45 T. Arai, Y. Ogawa, H. Sakuragi and K. Tokumaru, *Chem. Phys. Lett.*, 1992, **196**, 145–149.
- 46 F. Bejarano, I. J. Olavarria-Contreras, A. Droghetti, I. Rungger, A. Rudnev, D. Gutiérrez, M. Mas-Torrent, J. Veciana, H. S. J. Van Der Zant, C. Rovira, E. Burzurl and N. Crivillers, *J. Am. Chem. Soc.*, 2018, **140**, 1691–1696.
- 47 D. Fracasso, S. Kumar, P. Rudolf and R. C. Chiechi, *RSC Adv.*, 2014, **4**, 56026–56030.
- 48 S. Zhang, K. L. Chandra and C. B. Gorman, *J. Am. Chem. Soc.*, 2007, **129**, 4876–4877.
- 49 T. Zaba, A. Noworolska, C. M. Bowers, B. Breiten, G. M. Whitesides and P. Cyganik, *J. Am. Chem. Soc.*, 2014, **136**, 11918–11921.
- 50 P. Pla-Vilanova, A. C. Aragonès, S. Ciampi, F. Sanz, N. Darwish and I. Díez-Pérez, *Nanotechnology*, 2015, **26**, 381001.
- 51 M. M. Thuo, W. F. Reus, C. A. Nijhuis, J. R. Barber, C. Kim, M. D. Schulz and G. M. Whitesides, *J. Am. Chem. Soc.*, 2011, **133**, 2962–2975.
- 52 L. Yuan, C. Franco, N. Crivillers, M. Mas-Torrent, L. Cao, C. S. S. Sangeeth, C. Rovira, J. Veciana and C. A. Nijhuis, *Nat. Commun.*, 2016, **7**, 12066–12076.
- 53 M. Souto, V. Díez-Cabanes, L. Yuan, A. R. Kyvik, I. Ratera, C. A. Nijhuis, J. Cornil and J. Veciana, *Phys. Chem. Chem. Phys.*, 2018, **20**, 25638–25647.
- 54 C. A. Nijhuis, W. F. Reus and G. M. Whitesides, *J. Am. Chem. Soc.*, 2009, **131**, 17814–17827.
- 55 P. Song, L. Yuan, M. Roemer, L. Jiang and C. A. Nijhuis, *J. Am. Chem. Soc.*, 2016, **138**, 5769–5772.
- 56 S. Kumar, M. Merelli, W. Danowski, P. Rudolf, B. L. Feringa and R. C. Chiechi, *Adv. Mater.*, 2019, **31**, 1807831–1807836.
- 57 S. Kumar, J. T. Van Herpt, R. Y. N. Gengler, B. L. Feringa, P. Rudolf and R. C. Chiechi, *J. Am. Chem. Soc.*, 2016, **138**, 12519–12526.
- 58 C. A. Nijhuis, W. F. Reus, J. R. Barber, M. D. Dickey and G. M. Whitesides, *Nano Lett.*, 2010, **10**, 3611–3619.
- 59 J. M. Mativetsky, G. Pace, M. Elbing, M. A. Rampi, M. Mayor and P. Samorì, *J. Am. Chem. Soc.*, 2008, **130**, 9192–9193.
- 60 J. Chen, S. Gathiaka, Z. Wang and M. Thuo, *J. Phys. Chem. C*, 2017, **121**, 23931–23938.
- 61 A. Kovalchuk, T. Abu-Husein, D. Fracasso, D. A. Egger, E. Zojer, M. Zharnikov, A. Terfort and R. C. Chiechi, *Chem. Sci.*, 2016, **7**, 781–787.
- 62 A. Kovalchuk, D. A. Egger, T. Abu-Husein, E. Zojer, A. Terfort and R. C. Chiechi, *RSC Adv.*, 2016, **6**, 69479–69483.
- 63 V. Cantatore, G. Granucci, G. Rousseau, G. Padula and M. Persico, *J. Phys. Chem. Lett.*, 2016, **7**, 4027–4031.
- 64 A. Valentini, M. Nucci, L. M. Frutos and M. Marazzi, *ChemPhotoChem*, 2019, **3**, 925–932.
- 65 P. Farahani, M. Lundberg, R. Lindh and D. Roca-Sanjuán, *Phys. Chem. Chem. Phys.*, 2015, **17**, 18653–18664.
- 66 M. Marazzi, H. Gattuso, A. Giussani, H. Zhang, M. Navarrete-Miguel, C. Chipot, W. Cai, D. Roca-Sanjuán, F. Dehez and A. Monari, *J. Phys. Chem. Lett.*, 2019, **10**, 7133–7140.
- 67 R. González-Luque, G. Olaso-González, M. Merchán, P. B. Coto, L. Serrano-Andrés and M. Garavelli, *Int. J. Quantum Chem.*, 2011, **111**, 3431–3437.
- 68 B. G. Levine and T. J. Martínez, *Annu. Rev. Phys. Chem.*, 2006, **58**, 613–634.
- 69 It is noted that the B3LYP functional often suffers from an improper description of CT transitions; see ref. 76; the performance, however, depends sensitively on degree and nature of the CT character, see e.g. the example in ref. 77.
- 70 A. Dreuw and M. Head-Gordon, *Chem. Rev.*, 2005, **105**, 4009–4037.



- 71 B. Milián-Medina and J. Gierschner, *Org. Electron.*, 2012, **13**, 895–911.
- 72 J. Shi, M. A. Izquierdo, S. Oh, S. Y. Park, B. Milián-Medina, D. Roca-Sanjuán and J. Gierschner, *Org. Chem. Front.*, 2019, **6**, 948–1954.
- 73 J. González-Vázquez and L. González, *Chem. Phys.*, 2008, **349**, 287–295.
- 74 A. Nenov and R. De Vivie-Riedle, *J. Chem. Phys.*, 2012, **137**, 74101.
- 75 L. A. Estrada, A. Francés-Monerris, I. Schapiro, M. Olivucci and D. Roca-Sanjuán, *Phys. Chem. Chem. Phys.*, 2016, **18**, 32786–32795.
- 76 D. Roca-Sanjuán, M. Merchán, L. Serrano-Andrés and M. Rubio, *J. Chem. Phys.*, 2008, **129**, 095104–095114.

