

**Industry and Research in Chemistry: Specialized  
Topics in Theory and Practice**

Code: 42424  
ECTS Credits: 9

Degree	Type	Year	Semester
4313385 Industrial Chemistry and Introduction to Chemical Research	OB	0	1

## Contact

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## Teaching groups languages

You can check it through this [link](#). To consult the language you will need to enter the CODE of the subject. Please note that this information is provisional until 30 November 2023.

## Teachers

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## Prerequisites

none

## Objectives and Contextualisation

Industrial Chemistry and Introduction to Chemical Research is a mandatory module of the Masters Program in Industrial Chemistry and Introduction to Chemical Research. The objective of this course is to acquire new knowledge and abilities in fields related to Industrial Chemistry and Chemical Research: Regulations, Patents,

Experimental design, Job searching, Introduction to computational chemistry, NMR , Introduction to photochemistry, Resource optimization and environmental assessment of chemical processes, Risk and safety in chemical facilities, and Laboratory instrumental techniques and chemical analysis (including mass spectrometry, chromatography, microscopy, XRD, ICP and laboratory advanced techniques).

## Competences

- Correctly apply new information capture and organisation technologies to solve problems in professional activity.
- Correctly evaluate the risks and environmental and socio-economic impact associated to special chemical substances.
- Define specialised concepts, principles, theories and facts in the different areas of Chemistry.
- Design processes that imply the treatment or elimination of dangerous chemical products.
- Evaluate the human, economic, legal and ethical dimension of professional practice, as well as the environmental implications of one's work.
- Identify information in the scientific literature using the appropriate channels and integrating said information to approach and contextualise a research issue.
- Operate with advanced instrumentation for chemical evaluation and structural determination.
- Students should be able to integrate knowledge and face the complexity of making judgements from information which, being incomplete or limited, include reflections on the social and ethical responsibilities linked to the application of their knowledge and judgements
- Students should know how to communicate their conclusions, knowledge and final reasoning that they hold in front of specialist and non-specialist audiences clearly and unambiguously
- Use scientific terminology in the English language to defend experimental results in the context of the chemistry profession.

## Learning Outcomes

1. Apply advanced analytical and instrumental techniques in a chemistry laboratory.
2. Characterise materials and biomolecules.
3. Compare microscopy and spectroscopy techniques for applications of differing natures.
4. Correctly apply new information capture and organisation technologies to solve problems in professional activity.
5. Describe quality and patent regulations.
6. Design chemical experiments.
7. Design chemical processes that respect the environment.
8. Evaluate risks and security in chemical facilities and laboratories.
9. Evaluate the human, economic, legal and ethical dimension of professional practice, as well as the environmental implications of one's work.
10. Identify information in the scientific literature using the appropriate channels and integrating said information to approach and contextualise a research issue.
11. Know the environmental risks associated to special substances and chemical processes.
12. Students should be able to integrate knowledge and face the complexity of making judgements from information which, being incomplete or limited, include reflections on the social and ethical responsibilities linked to the application of their knowledge and judgements
13. Students should know how to communicate their conclusions, knowledge and final reasoning that they hold in front of specialist and non-specialist audiences clearly and unambiguously
14. Use scientific terminology in the English language to defend experimental results in the context of the chemistry profession.

## Content

## M1: Industry and Research in Chemistry: Specialized Topics in Theory and Practice

- Regulations.
- Patents.
- Experimental design.
- Job searching.
- Introduction to computational chemistry.
- NMR (theory + problem resolution + introductory practical course)
- Introduction to photochemistry.
- Resource optimization and environmental assessment of chemical processes.
- Risk and safety in chemical facilities.
- Laboratory instrumental techniques and chemical analysis.

Mass spectrometry, chromatography

Microscopy

XRD

ICP

Laboratory advanced techniques

## Methodology

Lectures / Individual work:

The student acquires the knowledge of the course by attending the lectures and complementing them with the individual work. The lectures may include problem solving (theoretically based or practical exercises) and seminars.

Laboratory work

Laboratory practices will be held to achieve specific and relevant competencies.

Annotation: Within the schedule set by the centre or degree programme, 15 minutes of one class will be reserved for students to evaluate their lecturers and their courses or modules through questionnaires.

## Activities

Title	Hours	ECTS	Learning Outcomes
Type: Directed			
Laboratory Practice	16	0.64	4, 1, 8, 3, 11, 6, 7, 10, 12, 13, 14, 9
Lectures	43	1.72	4, 1, 8, 2, 3, 11, 5, 6, 7, 10, 12, 14, 9
Type: Autonomous			
Individual work	146	5.84	4, 8, 2, 3, 11, 5, 6, 7, 10, 12, 13, 14, 9

## Assessment

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### Assessment

- Every professor decides the number and typology of evaluation activities: oral presentations, written exams, delivery of discussed articles, tests.
- The final mark of the module will be the sum of the marks of every professor multiplied by the percentage of his classes in the total teaching of the module.
- To pass a module, it is mandatory a mark over 5 in a 75% of all the activities in order to average with other marks of the professor and/or the module.
- There will be a period in January to repeat written exams with marks under 5. In the case of exams under 3,5, it will be mandatory to the student. In the case of exams between 3,5 and 5 it would be optional.
- The marks of other evaluations activities (i. e. oral presentations) will average with the rest of the marks of the professor/module independently of the value. There will be not option of repeating these evaluation activities.

**VERY IMPORTANT:** Partial or total plagiarising will immediately result in a FAIL (0) for the plagiarised exercise and the WHOLE subject. PLAGIARISING consists of copying text from unacknowledged sources -whether this is part of a sentence or a whole text - with the intention of passing it off as the student's own production. It includes cutting and pasting from internet sources, presented unmodified in the student's own text. Plagiarising is a **SERIOUS OFFENCE**. Students must respect authors' intellectual property, always identifying the sources they may use; they must also be responsible for the originality and authenticity of their own texts.

In the event of a student committing any irregularity that may lead to a significant variation in the grade awarded to an assessment activity, the student will be given a zero for this activity, regardless of any disciplinary process that may take place. In the event of several irregularities in assessment activities of the same subject, the student will be given a zero as the final grade for this subject.

**One single final assessment:** Students who have opted for the "One single final assessment" modality will have to take a final test consisting of an exam covering the entire theoretical content and problem-solving exercises of the subject. This test will be held on the period as the regular assessment students take the examinations. The student's final grade will be calculated as follows:  $\text{Subject Grade} = (\text{Final Test Grade} * 0.85 + \text{Laboratory Grade} * 0.15)$ . If the final grade does not reach 5, the student has another opportunity to pass the subject through a recovery exam, which will take place on a date set by the degree program coordination. In this recovery test, it will be possible to recover 85% of the grade corresponding to the theoretical part. The practical component is not recoverable. It is mandatory to pass the laboratory (minimum grade of 5.0).

## Assessment Activities

Title	Weighting	Hours	ECTS	Learning Outcomes
Presentations	20%	4	0.16	4, 8, 11, 6, 7, 10, 12, 13, 14, 9
Reports	40%	6	0.24	4, 1, 8, 2, 3, 11, 5, 6, 7, 10, 12, 13, 14, 9
Theoretical - practical tests	40%	10	0.4	4, 1, 8, 2, 3, 11, 5, 6, 7, 10, 12, 13, 14, 9

## Bibliography

## Patents

<http://www.ub.edu/centrepatents/es/>  
[http://www.oepm.es/es/propiedad\\_industrial/index.html](http://www.oepm.es/es/propiedad_industrial/index.html)  
<http://www.epo.org/law-practice.html>  
[http://e-courses.epo.org/wbts/htgaep\\_en/index.html](http://e-courses.epo.org/wbts/htgaep_en/index.html)

## Experimental design

Richard G. Brereton, "Applied Chemometrics for Scientists", 2007, chapter 2, Wiley Chichester.  
Rolf Carlson & Johan E. Carlson, "Design and optimization in organic synthesis", 2005, Series Data Handling in Science and Technology Vol. 24, Elsevier Amsterdam.  
Gareth A. Lewis; Didier Mathieu & Roger Phan-Tan-Luu, "Pharmaceutical experimental design", 1999, Marcel Dekker NY.

## Introduction to computational chemistry

C. J. Cramer, "Essentials of Computational Chemistry: Theories and Models", 2004, Wiley, 2nd edition.  
F. Jensen, "Introduction to Computational Chemistry", 2007, Wiley, 2nd edition.  
E. G. Lewars, "Computational Chemistry: Introduction to the Theory and Applications of Molecular and Quantum Mechanics", 2011, Springer, 2nd edition.  
D. C. Young, "Computational Chemistry: A Practical Guide for Applying Techniques to Real-World Problems", 2001, Wiley.

## Introduction to Photochemistry

Angelo Albini, "Photochemistry: Past, Present and Future", 2016, Springer-Verlag, Berlin Heidelberg.

## Resource optimization and environmental assessment of chemical processes

Xavier Domènech, "Química Verde", 2005, Editorial Rubes, ISBN 9788449701818.  
Risk Assessment and Sustainable Chemistry: <http://www.epa.gov/nrmrl/std/index.html>  
Life Cycle Assessment: <http://www.epa.gov/nrmrl/std/lca/resources.html>  
Donald Mackay, "Multimedia Environmental Models", 2001, Lewis Publishers, ISBN 1-56670-542-8.

## Risk and safety in chemical facilities

D.J. Knight, "EU Regulation of Chemicals: REACH", 2005, Rapra Review ReportN. 181, RapraTechnology Limited, Shawbury UK.

Regulation (EC) No 1907/2006 of the European Parliament and of the Council of 18 December 2006 concerning the Registration, Evaluation, Authorisation and Restriction of Chemicals (REACH), establishing a European Chemicals Agency.

## Laboratory instrumental techniques and chemical analysis

Thomas T. Tidwel, "Wilhelm Schlenk: The Man Behind the Flask", Angew. Chem. Int. ed. 2001, 40, 331-337.  
Duward F. Shriver, M. A. Drezdson, "The Manipulation of Air-Sensitive Compounds", 1986, J. Wiley and Sons: New York.

## Software

ChemBioDraw