

Chemical Kinetics

Code: 102404
ECTS Credits: 6

2025/2026

Degree	Type	Year
Chemical Engineering	OB	2

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Teachers

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Teaching groups languages

You can view this information at the [end](#) of this document.

Prerequisites

It is recommended to have obtained the fundamental knowledge of the subjects Química Inorgànica i de l'Equilibri i Bases de l'Enginyeria Química

Objectives and Contextualisation

The objectives of the Chemical Kinetics course are to acquire and consolidate the fundamental concepts of kinetic, stoichiometry and reaction mechanisms as well as homogenous kinetics and heterogeneous catalytic kinetic. However, it is intended to acquire the skills specified in this teaching guide.

Competences

Chemical Engineering

- Apply scientific method to systems in which chemical, physical or biological transformations are produced both on a microscopic and macroscopic scale.
- Develop personal work habits.
- Develop thinking habits.

- Understand and apply the basic principles on which chemical engineering is founded, and more precisely: balances of matter, energy and thermodynamic momentum, phase equilibrium and kinetic chemical equilibrium of the physical processes of matter, energy and momentum transfer, and kinetics of chemical reactions
- Work in a team.

Learning Outcomes

1. Analyse a scientific study of the kinetics of a chemical reaction.
2. Apply and identify basic concepts related with chemical engineering.
3. Apply the concepts of heterogeneous catalytic chemical kinetics.
4. Apply the concepts of homogenous chemical kinetics.
5. Develop independent learning strategies.
6. Develop scientific thinking.
7. Use criteria to determine the control stage of heterogeneous catalytic processes.
8. Work autonomously.
9. Work cooperatively.

Content

1. Introduction. Fundamental concepts
 - 1.1. Types of chemical reactions
 - 1.2. Balance, kinetic and stoichiometry
 - 1.3. Stoichiometric model
 - 1.4. Reaction measurement of progress
 - 1.5. Concept of reaction speed
 - 1.6. Mechanism of reaction and kinetic model
2. Homogeneous kinetics
 - 2.1. Reactions in homogeneous media
 - 2.2 Dependence of concentration and temperature
 - 2.3. Determination of reaction velocity and kinetic equation
 - 2.4. Simultaneous reactions
3. Heterogeneous catalytic kinetics
 - 3.1. Catalysts. Materials and properties.
 - 3.2. Adsorption
 - 3.3. Kinetic models for catalyzed reactions by solids
 - 3.4. External and internal transfer of subject matter
 - 3.5. Methods to determine the controlling resistance
 - 3.6. Disabling solid catalysts

Activities and Methodology

Title	Hours	ECTS	Learning Outcomes
Type: Directed			
master class	30	1.2	
Problem solving	15	0.6	
Type: Supervised			
Seminars	5.5	0.22	
Works	8	0.32	1
Type: Autonomous			
article searching	5	0.2	1
books and articles reading	4.5	0.18	1
Problem solving	35	1.4	
Study	32	1.28	1
tutorials	6	0.24	

The teaching methodology of this subject combines different elements: master classes and seminars; problem-based learning; analysis of scientific work; cooperative learning; peer evaluation; flipped classrooms. Communication with students will be via Virtual Campus (Moodle).

Annotation: Within the schedule set by the centre or degree programme, 15 minutes of one class will be reserved for students to evaluate their lecturers and their courses or modules through questionnaires.

Assessment

Continuous Assessment Activities

Title	Weighting	Hours	ECTS	Learning Outcomes
Seminars, works and evaluable tests	35%	3	0.12	1, 2, 9, 8
Written evaluation tests	65%	6	0.24	3, 4, 6, 5, 7

Process and evaluation activities programmed

The subject consists of the following evaluation activities:

1. Seminars and tests that can be evaluated

- First evaluable test PA1. Evaluable seminar, problem-based learning, problem solving in a collaborative way. 5% of the mark of the subject

- Second evaluative test PA2. Exercise on the design and applications of solid catalysts. Peer evaluation. 15% of the course grade.

- Third evaluative test PA3. Group work + Evaluated seminar. 15% of the course grade.

People repeating the subject will be able to choose to keep the grade of the evaluative tests from the previous year as a whole or repeat them all again.

2. Examinations

First exam on Topics 1 and 2. 30% of the course grade.

Second synthesis exam. 35% of the course grade.

Both exams will include a *question zero* about a basic concept that will have been covered and indicated in class beforehand. An incorrect answer to the zero question will result in immediate failure of the exam, and the student will have to take a makeup exam.

It is necessary to obtain a score of 4,5 out of 10 or higher in the weighted average of the exams to combine it with the evaluative tests and be eligible to pass the course. If this threshold is not reached, the grade for the subject will correspond to the weighted average of the exams.

Very important: Written submissions and exams must be presented with an adequate level of care and clarity; those that do not meet these standards will be penalized or, if the presentation is especially poor, may not be graded.

Programming of evaluation activities

The scheduling of the assessment activities will be given on the first day of the subject and will be made public through the Virtual Campus.

Recovery process

The recovery of the subject will be done through a synthesis examination of the syllabus corresponding to the entire subject, if they obtain a grade of less than 4.5 on the weighted average of the exams or less than 5 in the total of the subject.

Procedure for the review of qualifications

For each assessment activity, a place, date and time of revision will be indicated in which the student will be able to review the activity with the teacher. In this context, claims can be made about the activity note, which will be evaluated by the teachers responsible for the subject. If the student does not submit to this review, this activity will not be reviewed later.

Qualifications

Honor enrollments Granting an honorific matriculation qualification is a decision of the faculty responsible for the subject. The regulations of the UAB indicate that MH can only be awarded to students who have obtained a final grade of 9.00 or more. It can be granted up to 5% of MH of the total number of students enrolled.

A student will be considered non-evaluable (NA) if it has not been presented in a set of activities whose weight equals to a minimum of two thirds of the total grade of the subject.

Irregularities by the student, copy and plagiarism

Notwithstanding other disciplinary measures deemed appropriate, the irregularities committed by the student that can lead to a variation in the rating of an evaluation act will be graded with a zero. Therefore, copying, plagiarizing, cheating, copying, etc. In any of the assessment activities it will imply suspending it with a zero. Assessment activities qualified in this way and by this procedure will not be recoverable. If it is necessary to pass any of these assessment activities to pass the subject, this subject will be suspended directly, without opportunity to recover it in the same course.

RESTRICTED USE OF AI

Restricted use: For this course, the use of Artificial Intelligence (AI) technologies is allowed exclusively for tasks involving text correction or improvement, but not for content generation. Students must clearly indicate how they have used these technologies, specify the tools employed, and include a critical reflection on how

these influenced both the process and the final outcome of the activity. Lack of transparency in the use of AI in assessable activities will be considered academic dishonesty and may result in partial or total penalties to the activity grade, or more serious sanctions in severe cases.

Bibliography

Izquierdo, J.F., y col. "Cinética de las Reacciones Químicas". Ed. U.B., 2004.

Fogler, H.S. "Elements of Chemical Reaction Engineering", 4th ed., Prentice-Hall, 2005.

Levenspiel, O. "Chemical Reaction Engineering", 3rd. ed., John Wiley and Sons, 1999

On-line resources

Chemical Kinetics and Reaction Dynamics

By: Houston, Paul L. Courier Dover Publications.

https://app.knovel.com/web/toc.v/cid:kpCKRD0001/viewerType:toc/root_slug:chemical-kinetics-reaction?kpromot

Principles of Chemical Kinetics

By: House, J. E. Academic Press

<https://ebookcentral.proquest.com/lib/uab/reader.action?docID=535154&ppg=90>

An Introduction to Chemical Kinetics

By: Margaret Robson Wright. John Wiley & Sons Incorporated

<https://onlinelibrary.wiley.com/doi/book/10.1002/047009060X>

Software

MS Excel

Groups and Languages

Please note that this information is provisional until 30 November 2025. You can check it through this [link](#). To consult the language you will need to enter the CODE of the subject.

Name	Group	Language	Semester	Turn
(PAUL) Classroom practices	211	Catalan/Spanish	first semester	morning-mixed
(PAUL) Classroom practices	212	Catalan/Spanish	first semester	morning-mixed
(SEM) Seminars	211	Catalan/Spanish	first semester	morning-mixed
(SEM) Seminars	212	Catalan/Spanish	first semester	morning-mixed
(TE) Theory	21	Catalan/Spanish	first semester	morning-mixed