

Degree	Type	Year
Chemistry	FB	1

Contact

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Teachers

Gregorio Ujaque Perez

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Teaching groups languages

You can view this information at the [end](#) of this document.

Prerequisites

There are no official prerequisites. However, at the beginning of the course, students must know the fundamental concepts corresponding to the Chemistry courses (precipitation and redox).

Objectives and Contextualisation

The main objective of the subject is double. The first objective of this introductory course is to homogenize the level of knowledge of the students.

pre-university studies. Based on this knowledge, the second objective is to provide the student with the necessary

In particular and among other knowledge, the course must provide

security to

the student in complex stoichiometric calculations and the formulation and nomenclature of the most important

chemical compounds; qualitative knowledge of the structure of the atom and the types of bonds present in molecules; knowledge of the most important organic functional groups and the types of isomerism they show.

Learning Outcomes

1. CM01 (Competence) Interpret data obtained from experiments or models to propose solutions to problems in the field of general chemistry.
2. CM03 (Competence) Work autonomously in the field of chemistry, integrating knowledge and skills for problem solving, preparing laboratory protocols and delivering exercises and reports.
3. CM03 (Competence) Work autonomously in the field of chemistry, integrating knowledge and skills for problem solving, preparing laboratory protocols and delivering exercises and reports.
4. KM01 (Knowledge) Relate the structure of the atom, chemical bonding, intermolecular forces, and states of aggregation to the properties of matter.
5. KM02 (Knowledge) Identify concepts, principles and theories in the field of thermochemistry, homogeneous and heterogeneous equilibria, chemical kinetics and electrochemistry.
6. SM01 (Skill) Accurately use the terminology of chemical compounds, chemical equations and magnitudes of chemistry.
7. SM02 (Skill) Determine the properties of elements and simple molecules by applying Lewis' theories, valence bond theory, and molecular orbital theory.
8. SM03 (Skill) Correctly perform calculations on simple chemical reactions from a thermodynamic and perspective to predict their evolution.

Content

PART I. Matter, compounds and chemical reactions

Chapter 1. Matter and chemical compounds

Chapter 2. Introduction to chemical reactions

Chapter 3. Gases

PART II. Atomic structure and bonding

Chapter 4. Atomic structure

Chapter 5. The periodic table

Chapter 6. Chemical bonding

Chapter 7. Bonding in solids and liquids

Unless the requirements enforced by the health authorities demand a prioritization or reduction of these contents

Activities and Methodology

Title	Hours	ECTS	Learning Outcomes
Type: Directed			
Exercice lessons	20	0.8	
Study	106	4.24	
Theoretical lessons	48	1.92	

The course Fonaments de Química I consists of two types of supervised activities, the theoretical sessions and the excercise sessions, which are distributed throughout the course in an alternating way. Theoretical sessions. Through the teacher's expositions the student must

of the subject and complement it with his/her personal study with the help of the materials that the teacher has provided through the Campus Virtual/Teams and the recommended bibliography. The theoretical sessions will be open to the participation of the students, who

the questions and clarifications that they deem necessary. Excercise sessions. The objective of this supervised activity is to solve problems that have been previously raised to students through the Campus Virtual and were asked to be resolved previously, in group or individually. We aim to stimulate the participation of the students in solving problems, taking advantage of it to consolidate the knowledge acquired during the theoretical sessions and during their personal study.

The proposed teaching methodology may experience some modifications depending on the restrictions to face-to-face teaching.

Annotation: Within the schedule set by the centre or degree programme, 15 minutes of one class will be reserved for students to evaluate their lecturers and their courses or modules through questionnaires.

Assessment

Continous Assessment Activities

Title	Weighting	Hours	ECTS	Learning Outcomes
Continuous work	20	20	0.8	CM01, CM03, KM01, KM02, SM01, SM02, SM03
Exams	80	6	0.24	CM01, CM03, KM01, KM02, SM01, SM02, SM03

Student evaluation will have two options: continuous evaluation and single evaluation.

OPTION A: Continuous Evaluation (this is the default option)

It will be done through several evaluation tests:

- Written exams:

There will be two partial exams throughout the course, one in the middle and the other at the end of the semester. Each of these exams will have a weight of 40% on the final mark (with a total of 80%). To be able to average the grade for each exam the mark must be ≥ 4.0 .

- Continued work:

Student evidences will be collected throughout the course (problems solved, individually or in groups, self-assessments on the Virtual Campus, short tests in class, etc.). These activities cannot be recovered unless the student provides a greater justification with the corresponding official documentation. This activity will have a weight of 20% of the grade.

Subject grade = Average grade of the partial tests (80%) + continuous work grade (20%)

Option B: Single Evaluation

Students who have accepted the single assessment modality will have to take a final test which will consist of an examination of the entire theoretical syllabus and problems of the subject. This test will be carried out on the day on which the students of the continuous assessment take the second part exam. The student's qualification will be:

Grade of the subject = Grade of the final test

Both for option A and B, the subject is passed with a 5. If the final grade does not reach 5, the student has another opportunity to pass the subject through the make-up exam that will be held on the date set by the degree coordination.

Not assessable:

The subject will be graded as "Not assessable" when: a) the student has not taken any of the partial and the final exams of the subject, or b) has delivered less than three activities of the Continued evaluation. If you have opted for the single assessment if you do not take the final test.

Bibliography

Textbook

QUIMICA GENERAL: PRINCIPIOS Y APLICACIONES MODERNAS, R. H. Petrucci, F. G. Herring, J.D. Madura i C. Bissonnette , Pearson Educación SA, 10^a edició, Madrid 2011 (ISBN: 978-84-8322-680-3).

http://www.ingebook.com.are.uab.cat/ib/NPcd/IB_BooksVis?cod_primaria=1000187&codigo_libro=1262

Other useful books:

CHEMISTRY: A MOLECULAR APPROACH, Nivaldo Tro, Pearson, 5th Ed. 2020

PRINCIPIOS DE QUÍMICA, P. Atkins i L. Jones, Médica Panamericana, 3^a edició, 2006.

QUÍMICA, R. Chang, McGraw-Hill, 9^a edició, 2010.

PRINCIPIOS DE FÍSICO-QUÍMICA, Ira N. Levine, McGraw-Hill 6^a edició, 2014

INTRODUCCIÓ A LA NOMECLATURA QUÍMICA INORGÀNICA I ORGÀNICA, J. Sales i J. Vilarrasa, Reverté, 5^a edició, 2003.

INTRODUCCIÓN A LA NOMENCLATURA DE LAS SUSTANCIAS QUÍMICAS, W. R. Peterson, Reverté, 2010.

Nomenclatura de Química:

Brief guides to Nomenclature IUPAC:

<https://iupac.org/what-we-do/nomenclature/brief-guides/>

Software

The course does not imply the mandatory use of any specific software.

Groups and Languages

Please note that this information is provisional until 30 November 2025. You can check it through this [link](#). To consult the language you will need to enter the CODE of the subject.

Name	Group	Language	Semester	Turn
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(PAUL) Classroom practices	1	Catalan	first semester	morning-mixed
(PAUL) Classroom practices	2	Catalan	first semester	morning-mixed
(PAUL) Classroom practices	3	Catalan	first semester	afternoon
(PAUL) Classroom practices	4	Catalan	first semester	afternoon
(TE) Theory	1	Catalan	first semester	morning-mixed
(TE) Theory	2	Catalan	first semester	afternoon